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CHAPTER 09 STOICHIOMETRY NOTES (ANSWERS)

Unit 5: Chemical Equations and Reactions & Stoichiometry Chemistry Chapter 9: Stoichiometry 9.1: Calculating Quantities in Reactions Avogadro's Number: - a group of (6.022 10²³) molecules = 1 mole Stoichiometry: - the calculation of quantities in a chemical reaction. - the coefficients of various reactants and or products form mole ratios. - these mole ratios hold for moles, molecules or atoms. Mole Rat...

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M:\USERS\H\HENZEL, SHERMAN\CHE100\LECTURE\STOICHIOMETRY PROBLEM ANSWERS.WPD

CHE100 Stoichiometry Problems Instructor: S. Henzel $2\text{C}_3\text{H}_6\text{O}_2 + 7\text{O}_2 \rightarrow 6\text{CO}_2 + 6\text{H}_2\text{O}$ 1. How many moles of CO_2 are produced when 8.652 moles of $\text{C}_3\text{H}_6\text{O}_2$ react? $8.652\text{ mol C}_3\text{H}_6\text{O}_2 \times \frac{6\text{ mol CO}_2}{2\text{ mol C}_3\text{H}_6\text{O}_2} = 25.96\text{ mol CO}_2$ 2. How many moles of H_2O are produced when 50.75 g of O_2 react? $50.75\text{ g O}_2 \times \frac{1\text{ mol O}_2}{32\text{ g O}_2} \times \frac{6\text{ mol H}_2\text{O}}{7\text{ mol O}_2} = 1.359\text{ mol H}_2\text{O}$ 3. How many moles of O_2 are requir...

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