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CHAPTER 10

Getting the other gas laws If temperature is constant If volume is constant you you get Boyles Law get Gay-Lussacs Law $P_1 V_1 = P_2 V_2$ $P_1 P_2 T_1 = T_2$ If pressure is constant you get Charles Law If all are variable, you get the Combined Gas Law. $V_1 V_2 T_1 = T_2$ $P_1 V_1 P_2 V_2 T_1 = T_2$ Boyles Law: Inverse relationship between Pressure and Volume. $P_1 V_1 = P_2 V_2$ P V Combined Gas Law: Gay-Lussacs Law: $P_1 V_1 P_2 V_2 \dots$

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MICROSOFT WORD - COMBINED GAS LAWS-NOTES

Combined Gas Laws $PV = \text{constant}$ $PV = \text{constant}$ (Boyles Law) T (combined gas law) $V = \text{constant}$ (Charles Law) $T P = \text{constant}$ TALL gas law problems must be work using absolute temperature, kelvin (K).Memory Work: $760 \text{ mmHg} = 760 \text{ torr} = 1 \text{ atm} = 1.013 \times 10^5 \text{ Pa}$ and $P_1 V_1 = P_2 V_2$ $T_1 T_2$ or $P_1 V_1 T_2 = P_2 V_2 T_1$ and STP (standard temperature and pressure) $0 \text{ }^\circ\text{C}$ and 1 atm (273 K) (760 torr) V Volume amount of gas cons...

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